

catena-Poly[[[bis[2,2'-(propane-1,3-diyl-dithio)bis(1,3,4-thiadiazole)- κ N⁴]-copper(II)]-bis[μ -2,2'-(propane-1,3-diyl-dithio)bis(1,3,4-thiadiazole)- κ^2 N⁴:N^{4'}]] bis(perchlorate)]

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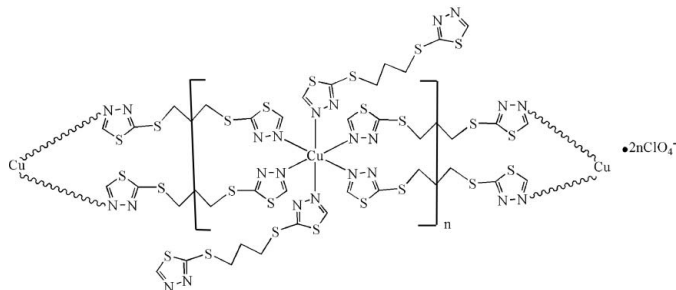
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Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.040; wR factor = 0.110; data-to-parameter ratio = 15.1.

In the title compound, $\{[\text{Cu}(\text{C}_7\text{H}_8\text{N}_4\text{S}_4)_4](\text{ClO}_4)_2\}_n$, the Cu^{II} atom, occupying a crystallographic inversion centre, is six-coordinated by six N atoms of three symmetry-related 2,2'-(propane-1,3-diyl-dithio)bis(1,3,4-thiadiazole) (*L*) ligands in a slightly distorted octahedral geometry. The ligand *L* adopts two kinds of coordination modes in the crystal structure; one is a monodentate coordination mode and serves to complete the octahedral coordination of the Cu atom and the other is an *N*:*N'*-bidentate bridging mode in a *trans* configuration, bridging Cu atoms *via* translation symmetry along the *b* axis into a chain structure. The perchlorate ions serve as acceptors for intermolecular C—H...O hydrogen bonds, which link the chains into a three-dimensional network.

Related literature

For Cu—N bonds see, for example: Huang *et al.* (2009); Wang *et al.* (2008).



Experimental

Crystal data

$[\text{Cu}(\text{C}_7\text{H}_8\text{N}_4\text{S}_4)_4](\text{ClO}_4)_2$	$\gamma = 76.675 (3)^\circ$
$M_r = 1368.10$	$V = 1312.3 (6) \text{ \AA}^3$
Triclinic, $P\bar{1}$	$Z = 1$
$a = 10.321 (3) \text{ \AA}$	Mo $K\alpha$ radiation
$b = 11.122 (3) \text{ \AA}$	$\mu = 1.22 \text{ mm}^{-1}$
$c = 12.908 (4) \text{ \AA}$	$T = 294 \text{ K}$
$\alpha = 67.213 (3)^\circ$	$0.39 \times 0.28 \times 0.24 \text{ mm}$
$\beta = 76.602 (3)^\circ$	

Data collection

Bruker SMART CCD area-detector diffractometer	9833 measured reflections
Absorption correction: multi-scan (<i>SADABS</i> ; Bruker, 1997)	4857 independent reflections
$T_{\min} = 0.646$, $T_{\max} = 0.756$	4081 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.017$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.040$	322 parameters
$wR(F^2) = 0.110$	H-atom parameters constrained
$S = 1.03$	$\Delta\rho_{\text{max}} = 1.05 \text{ e \AA}^{-3}$
4857 reflections	$\Delta\rho_{\text{min}} = -0.50 \text{ e \AA}^{-3}$

Table 1

Selected bond lengths (Å).

Cu1—N1	2.021 (2)	Cu1—N4 ⁱ	2.445 (3)
Cu1—N5	2.053 (2)		

Symmetry code: (i) $x, y - 1, z$.

Table 2

Hydrogen-bond geometry (Å, °).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
C3—H3B...O3 ⁱⁱⁱ	0.97	2.47	3.357 (6)	153
C7—H7...O3 ⁱⁱⁱ	0.93	2.51	3.172 (6)	128
C8—H8...O4 ⁱ	0.93	2.47	3.010 (6)	117
C10—H10A...O2 ^{iv}	0.97	2.50	3.423 (7)	159
C14—H14...O1 ^v	0.93	2.51	3.419 (7)	167

Symmetry codes: (i) $x, y - 1, z$; (ii) $x - 1, y, z$; (iii) $x - 1, y + 1, z$; (iv) $-x + 1, -y, -z + 2$; (v) $-x + 1, -y + 1, -z + 2$.

Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINTE* (Bruker, 1997); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SI2158).

References

- Bruker (1997). *SMART, SAINT and SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Huang, H.-M., Ju, F.-Y., Wang, J.-G. & Qin, J.-H. (2009). *Acta Cryst.* **E65**, m80–m81.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Wang, J. G., Qin, J. H., Hu, P. Z. & Zhao, B. T. (2008). *Z. Kristallogr. New Cryst. Struct.* **223**, 225–227.

supplementary materials

Acta Cryst. (2009). E65, m349-m350 [doi:10.1107/S1600536809006722]

***catena*-Poly[[[bis[2,2'-(propane-1,3-diylthio)bis(1,3,4-thiadiazole)- κ N⁴]]copper(II)]-bis[μ -2,2'-(propane-1,3-diylthio)bis(1,3,4-thiadiazole)- κ ²N⁴:N^{4'}]] bis(perchlorate)]**

J.-H. Qin, J.-G. Wang and P.-Z. Hu

Comment

The asymmetric structure unit of the title compound consists of a half Cu(II) atom, two [1,3-propanediylbis(thio)]bis[1,3,4-thiadiazole] ligands L, and one perchlorate ion. As depicted in Fig. 1, the Cu(II) atom is coordinated by six N atoms from six ligands L in a slightly distorted octahedral geometry of the central atom. All six Cu—N bond distances are within the range expected for such coordination bonds (Tab. 1) (Huang *et al.*, 2009; Wang *et al.*, 2008). The ligand L adopts two kinds of coordination modes in the crystal structure. One *N,N*-bidentate bridging mode in trans configuration for bridging the copper atom into a one-dimensional chain, with the bridged Cu—Cu distance of 11.122 (3) Å (Fig. 2). The centroid separation and dihedral angle of thiadiazole rings are 9.131 (2) Å and 74.09 (8) °, respectively. The other thiadiazole ligands adopt monodentate coordination mode and serve to complete the octahedral coordination sphere of the copper atom. The corresponding centroid separation and dihedral angle are 8.1499 (16) Å and 65.04 (12) °, respectively. The region between the chains is taken up by uncoordinated perchlorate ions. The perchlorate ions serve as acceptor for C—H···O hydrogen-bonds, which link the chains into a three-dimensional network (Tab. 2. & Fig. 3).

Experimental

The reaction of [1,3-propanediylbis(thio)]bis[1,3,4-thiadiazole] (0.4 mmol) with Cu(ClO₄)₂ (0.1 mmol) in MeOH(10 ml) for a few minutes afforded a light blue solid, which was filtered, washed with acetone, and dried on air. The single crystals suitable for X-ray analysis were obtained by slow diffusion of Et₂O into the acetonitrile solution of the solid.

Refinement

All hydrogen atoms were positioned geometrically and treated as riding, with C—H = 0.93 Å (CH) and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$, with C—H = 0.97 Å (CH₂) and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Figures

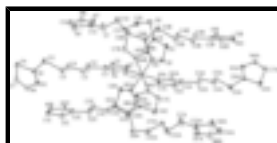


Fig. 1. A view of the local coordination of the Cu(II) cation in the title compound. Displacement ellipsoids are drawn at the 30% probability level. The H atoms and perchlorate ion were omitted for clarity. Symmetry codes: (A) $-x + 1, -y, -z + 1$; (B) $x, y - 1, z$; (C) $-x + 1, -y + 1, -z + 1$.



Fig. 2. A view of the polymeric chain in the title compound.

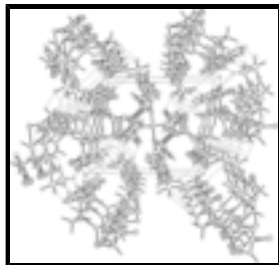


Fig. 3. A view of the compound packing down the *b* axis.

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Crystal data

[Cu(C₇H₈N₄S₄)₄](ClO₄)₂

$M_r = 1368.10$

Triclinic, $P\bar{1}$

Hall symbol: -P 1

$a = 10.321 (3) \text{ \AA}$

$b = 11.122 (3) \text{ \AA}$

$c = 12.908 (4) \text{ \AA}$

$\alpha = 67.213 (3)^\circ$

$\beta = 76.602 (3)^\circ$

$\gamma = 76.675 (3)^\circ$

$V = 1312.3 (6) \text{ \AA}^3$

$Z = 1$

$F_{000} = 695$

$D_x = 1.731 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation

$\lambda = 0.71073 \text{ \AA}$

Cell parameters from 4352 reflections

$\theta = 2.5\text{--}28.1^\circ$

$\mu = 1.22 \text{ mm}^{-1}$

$T = 294 \text{ K}$

Block, blue

$0.39 \times 0.28 \times 0.24 \text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 294 \text{ K}$

φ and ω scans

Absorption correction: multi-scan (SADABS; Bruker, 1997)

$T_{\min} = 0.646$, $T_{\max} = 0.756$

9833 measured reflections

4857 independent reflections

4081 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.017$

$\theta_{\max} = 25.5^\circ$

$\theta_{\min} = 2.5^\circ$

$h = -12 \rightarrow 12$

$k = -13 \rightarrow 13$

$l = -15 \rightarrow 15$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.040$

$wR(F^2) = 0.110$

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0533P)^2 + 1.507P]$

$S = 1.03$

4857 reflections

322 parameters

Primary atom site location: structure-invariant direct methods

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 1.05 \text{ e } \text{\AA}^{-3}$

$\Delta\rho_{\min} = -0.50 \text{ e } \text{\AA}^{-3}$

Extinction correction: none

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	0.5000	0.0000	0.5000	0.02910 (14)
Cl1	0.89564 (9)	0.39343 (8)	0.71273 (7)	0.0469 (2)
S1	0.60750 (10)	0.13642 (9)	0.74944 (8)	0.0494 (2)
S2	0.42947 (13)	0.39702 (10)	0.69426 (11)	0.0728 (4)
S3	-0.00185 (9)	0.78315 (9)	0.62689 (9)	0.0516 (2)
S4	0.02700 (10)	1.01392 (10)	0.68458 (9)	0.0569 (3)
S5	0.58105 (10)	-0.40131 (8)	0.77434 (7)	0.0458 (2)
S6	0.33608 (12)	-0.34525 (10)	0.94441 (8)	0.0615 (3)
S7	0.15521 (18)	0.11517 (12)	0.90859 (13)	0.0951 (5)
S8	0.11321 (15)	0.35074 (11)	0.97930 (11)	0.0777 (4)
O1	0.8989 (5)	0.4074 (5)	0.8153 (4)	0.1235 (17)
O2	0.8069 (6)	0.3095 (6)	0.7325 (4)	0.156 (2)
O3	1.0235 (4)	0.3459 (4)	0.6637 (4)	0.1009 (12)
O4	0.8505 (4)	0.5164 (4)	0.6350 (4)	0.1255 (17)
N1	0.5187 (2)	0.0858 (2)	0.6070 (2)	0.0307 (5)
N2	0.4504 (3)	0.2109 (2)	0.6008 (2)	0.0347 (6)
N3	0.2075 (2)	0.9204 (2)	0.5508 (2)	0.0371 (6)
N4	0.2550 (3)	1.0248 (2)	0.5561 (2)	0.0376 (6)
N5	0.5102 (3)	-0.1771 (2)	0.6333 (2)	0.0336 (5)
N6	0.4134 (3)	-0.1871 (2)	0.7283 (2)	0.0399 (6)
N7	0.2510 (4)	0.1319 (3)	1.0804 (3)	0.0609 (9)
N8	0.2522 (4)	0.2173 (4)	1.1350 (3)	0.0696 (10)
C1	0.6013 (3)	0.0364 (3)	0.6800 (3)	0.0366 (7)
H1	0.6529	-0.0470	0.6933	0.044*
C2	0.4879 (3)	0.2496 (3)	0.6710 (3)	0.0402 (7)
C3	0.3046 (4)	0.4799 (3)	0.6013 (3)	0.0510 (9)
H3A	0.3478	0.5122	0.5226	0.061*
H3B	0.2451	0.4200	0.6088	0.061*
C4	0.2261 (4)	0.5940 (3)	0.6378 (4)	0.0560 (10)
H4A	0.2884	0.6480	0.6366	0.067*
H4B	0.1801	0.5594	0.7154	0.067*
C5	0.1235 (4)	0.6790 (3)	0.5613 (3)	0.0466 (8)
H5A	0.0789	0.6227	0.5442	0.056*
H5B	0.1696	0.7341	0.4901	0.056*
C6	0.0906 (3)	0.9029 (3)	0.6139 (3)	0.0376 (7)
C7	0.1716 (3)	1.0817 (3)	0.6212 (3)	0.0462 (8)
H7	0.1887	1.1538	0.6326	0.055*
C8	0.6026 (3)	-0.2808 (3)	0.6447 (3)	0.0381 (7)

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H8	0.6739	-0.2882	0.5873	0.046*
C9	0.4384 (3)	-0.3003 (3)	0.8092 (3)	0.0418 (7)
C10	0.3970 (5)	-0.2527 (4)	1.0086 (3)	0.0651 (11)
H10A	0.3613	-0.2815	1.0893	0.078*
H10B	0.4945	-0.2750	1.0011	0.078*
C11	0.3601 (5)	-0.1024 (4)	0.9580 (3)	0.0645 (11)
H11A	0.4071	-0.0620	0.9898	0.077*
H11B	0.3898	-0.0730	0.8764	0.077*
C12	0.2129 (5)	-0.0586 (4)	0.9817 (4)	0.0760 (13)
H12A	0.1868	-0.0778	1.0630	0.091*
H12B	0.1665	-0.1103	0.9608	0.091*
C13	0.1830 (4)	0.1885 (4)	0.9971 (3)	0.0544 (9)
C14	0.1852 (4)	0.3311 (4)	1.0921 (4)	0.0614 (10)
H14	0.1762	0.3984	1.1203	0.074*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cu1	0.0416 (3)	0.0211 (2)	0.0251 (2)	0.00102 (19)	-0.0102 (2)	-0.00926 (19)
Cl1	0.0481 (5)	0.0429 (4)	0.0516 (5)	-0.0084 (4)	-0.0051 (4)	-0.0196 (4)
S1	0.0611 (6)	0.0441 (5)	0.0574 (5)	0.0031 (4)	-0.0342 (4)	-0.0256 (4)
S2	0.0904 (8)	0.0534 (6)	0.1059 (9)	0.0226 (5)	-0.0577 (7)	-0.0569 (6)
S3	0.0360 (4)	0.0434 (5)	0.0709 (6)	-0.0081 (4)	-0.0010 (4)	-0.0186 (4)
S4	0.0479 (5)	0.0579 (6)	0.0659 (6)	-0.0016 (4)	0.0071 (4)	-0.0358 (5)
S5	0.0629 (5)	0.0276 (4)	0.0415 (4)	0.0035 (4)	-0.0184 (4)	-0.0068 (3)
S6	0.0777 (7)	0.0484 (5)	0.0432 (5)	-0.0163 (5)	0.0037 (5)	-0.0036 (4)
S7	0.1475 (13)	0.0572 (7)	0.0988 (10)	0.0182 (8)	-0.0737 (10)	-0.0357 (7)
S8	0.1132 (10)	0.0470 (6)	0.0778 (8)	0.0187 (6)	-0.0536 (7)	-0.0231 (5)
O1	0.136 (4)	0.153 (4)	0.130 (3)	0.027 (3)	-0.063 (3)	-0.107 (3)
O2	0.191 (5)	0.200 (5)	0.137 (4)	-0.153 (5)	0.053 (4)	-0.098 (4)
O3	0.072 (2)	0.086 (2)	0.125 (3)	0.0190 (18)	0.004 (2)	-0.045 (2)
O4	0.104 (3)	0.091 (3)	0.107 (3)	0.038 (2)	-0.002 (2)	0.005 (2)
N1	0.0351 (13)	0.0263 (12)	0.0310 (12)	0.0002 (10)	-0.0080 (10)	-0.0118 (10)
N2	0.0408 (14)	0.0289 (12)	0.0397 (14)	0.0025 (10)	-0.0136 (11)	-0.0183 (11)
N3	0.0336 (13)	0.0342 (13)	0.0441 (15)	-0.0026 (10)	-0.0057 (11)	-0.0162 (11)
N4	0.0381 (14)	0.0330 (13)	0.0433 (15)	-0.0018 (11)	-0.0092 (11)	-0.0155 (12)
N5	0.0445 (14)	0.0261 (12)	0.0295 (13)	-0.0005 (10)	-0.0091 (11)	-0.0103 (10)
N6	0.0473 (15)	0.0330 (13)	0.0338 (14)	-0.0022 (11)	-0.0054 (11)	-0.0085 (11)
N7	0.071 (2)	0.0506 (18)	0.0563 (19)	0.0094 (16)	-0.0230 (17)	-0.0175 (16)
N8	0.086 (3)	0.064 (2)	0.064 (2)	0.0160 (19)	-0.0347 (19)	-0.0300 (18)
C1	0.0423 (17)	0.0323 (15)	0.0371 (16)	0.0015 (13)	-0.0144 (13)	-0.0140 (13)
C2	0.0447 (18)	0.0357 (16)	0.0486 (18)	0.0006 (14)	-0.0164 (14)	-0.0228 (15)
C3	0.053 (2)	0.0424 (19)	0.064 (2)	0.0020 (16)	-0.0183 (18)	-0.0267 (18)
C4	0.066 (2)	0.0387 (19)	0.070 (3)	0.0057 (17)	-0.024 (2)	-0.0267 (18)
C5	0.051 (2)	0.0367 (17)	0.056 (2)	-0.0069 (15)	-0.0110 (16)	-0.0183 (16)
C6	0.0355 (16)	0.0337 (16)	0.0398 (17)	0.0008 (13)	-0.0064 (13)	-0.0123 (13)
C7	0.050 (2)	0.0391 (18)	0.054 (2)	0.0009 (15)	-0.0135 (16)	-0.0217 (16)
C8	0.0463 (18)	0.0332 (16)	0.0357 (16)	0.0015 (13)	-0.0133 (13)	-0.0135 (13)

C9	0.054 (2)	0.0306 (16)	0.0377 (17)	-0.0075 (14)	-0.0097 (14)	-0.0073 (13)
C10	0.078 (3)	0.064 (3)	0.046 (2)	0.000 (2)	-0.0054 (19)	-0.0201 (19)
C11	0.084 (3)	0.064 (3)	0.049 (2)	-0.021 (2)	-0.001 (2)	-0.024 (2)
C12	0.088 (3)	0.052 (2)	0.087 (3)	0.001 (2)	-0.020 (3)	-0.027 (2)
C13	0.065 (2)	0.0404 (19)	0.053 (2)	0.0016 (17)	-0.0184 (18)	-0.0126 (17)
C14	0.071 (3)	0.054 (2)	0.062 (2)	0.004 (2)	-0.020 (2)	-0.025 (2)

Geometric parameters (Å, °)

Cu1—N1 ⁱ	2.021 (2)	N3—N4	1.390 (4)
Cu1—N1	2.021 (2)	N4—C7	1.293 (4)
Cu1—N5 ⁱ	2.053 (2)	N4—Cu1 ^{iv}	2.445 (3)
Cu1—N5	2.053 (2)	N5—C8	1.299 (4)
Cu1—N4 ⁱⁱ	2.445 (3)	N5—N6	1.375 (3)
Cu1—N4 ⁱⁱⁱ	2.445 (3)	N6—C9	1.304 (4)
Cl1—O2	1.370 (4)	N7—C13	1.291 (5)
Cl1—O4	1.398 (4)	N7—N8	1.387 (5)
Cl1—O3	1.404 (3)	N8—C14	1.271 (5)
Cl1—O1	1.400 (4)	C1—H1	0.9300
S1—C1	1.696 (3)	C3—C4	1.517 (5)
S1—C2	1.735 (3)	C3—H3A	0.9700
S2—C2	1.727 (3)	C3—H3B	0.9700
S2—C3	1.806 (4)	C4—C5	1.514 (5)
S3—C6	1.746 (3)	C4—H4A	0.9700
S3—C5	1.815 (4)	C4—H4B	0.9700
S4—C7	1.714 (4)	C5—H5A	0.9700
S4—C6	1.735 (3)	C5—H5B	0.9700
S5—C8	1.695 (3)	C7—H7	0.9300
S5—C9	1.719 (3)	C8—H8	0.9300
S6—C9	1.767 (3)	C10—C11	1.530 (6)
S6—C10	1.829 (5)	C10—H10A	0.9700
S7—C13	1.742 (4)	C10—H10B	0.9700
S7—C12	1.816 (5)	C11—C12	1.483 (6)
S8—C14	1.703 (4)	C11—H11A	0.9700
S8—C13	1.727 (4)	C11—H11B	0.9700
N1—C1	1.293 (4)	C12—H12A	0.9700
N1—N2	1.388 (3)	C12—H12B	0.9700
N2—C2	1.302 (4)	C14—H14	0.9300
N3—C6	1.299 (4)		
N1 ⁱ —Cu1—N1	180.0	C4—C3—H3B	110.5
N1 ⁱ —Cu1—N5 ⁱ	88.01 (9)	S2—C3—H3B	110.5
N1—Cu1—N5 ⁱ	91.99 (9)	H3A—C3—H3B	108.7
N1 ⁱ —Cu1—N5	91.99 (9)	C3—C4—C5	112.1 (3)
N1—Cu1—N5	88.01 (9)	C3—C4—H4A	109.2
N5 ⁱ —Cu1—N5	180.0	C5—C4—H4A	109.2
N1 ⁱ —Cu1—N4 ⁱⁱ	91.50 (9)	C3—C4—H4B	109.2
N1—Cu1—N4 ⁱⁱ	88.50 (9)	C5—C4—H4B	109.2

supplementary materials

N5 ⁱ —Cu1—N4 ⁱⁱ	87.34 (9)	H4A—C4—H4B	107.9
N5—Cu1—N4 ⁱⁱ	92.66 (9)	C4—C5—S3	111.8 (3)
N1 ⁱ —Cu1—N4 ⁱⁱⁱ	88.50 (9)	C4—C5—H5A	109.3
N1—Cu1—N4 ⁱⁱⁱ	91.50 (9)	S3—C5—H5A	109.3
N5 ⁱ —Cu1—N4 ⁱⁱⁱ	92.66 (9)	C4—C5—H5B	109.3
N5—Cu1—N4 ⁱⁱⁱ	87.34 (9)	S3—C5—H5B	109.3
N4 ⁱⁱ —Cu1—N4 ⁱⁱⁱ	180.0	H5A—C5—H5B	107.9
O2—C11—O4	108.6 (4)	N3—C6—S4	114.1 (2)
O2—C11—O3	110.1 (3)	N3—C6—S3	125.0 (2)
O4—C11—O3	107.8 (2)	S4—C6—S3	120.83 (18)
O2—C11—O1	108.9 (3)	N4—C7—S4	114.8 (3)
O4—C11—O1	109.3 (3)	N4—C7—H7	122.6
O3—C11—O1	112.2 (3)	S4—C7—H7	122.6
C1—S1—C2	86.96 (14)	N5—C8—S5	113.6 (2)
C2—S2—C3	103.64 (16)	N5—C8—H8	123.2
C6—S3—C5	101.32 (16)	S5—C8—H8	123.2
C7—S4—C6	86.73 (16)	N6—C9—S5	114.4 (2)
C8—S5—C9	87.49 (15)	N6—C9—S6	122.9 (3)
C9—S6—C10	99.43 (18)	S5—C9—S6	122.70 (18)
C13—S7—C12	102.4 (2)	C11—C10—S6	115.4 (3)
C14—S8—C13	86.78 (19)	C11—C10—H10A	108.4
C1—N1—N2	113.7 (2)	S6—C10—H10A	108.4
C1—N1—Cu1	124.3 (2)	C11—C10—H10B	108.4
N2—N1—Cu1	121.73 (17)	S6—C10—H10B	108.4
C2—N2—N1	110.3 (2)	H10A—C10—H10B	107.5
C6—N3—N4	111.9 (2)	C12—C11—C10	111.8 (4)
C7—N4—N3	112.5 (3)	C12—C11—H11A	109.2
C7—N4—Cu1 ^{iv}	133.3 (2)	C10—C11—H11A	109.2
N3—N4—Cu1 ^{iv}	109.59 (17)	C12—C11—H11B	109.2
C8—N5—N6	113.8 (2)	C10—C11—H11B	109.2
C8—N5—Cu1	128.7 (2)	H11A—C11—H11B	107.9
N6—N5—Cu1	117.31 (18)	C11—C12—S7	115.4 (4)
C9—N6—N5	110.7 (3)	C11—C12—H12A	108.4
C13—N7—N8	111.7 (3)	S7—C12—H12A	108.4
C14—N8—N7	112.8 (3)	C11—C12—H12B	108.4
N1—C1—S1	114.4 (2)	S7—C12—H12B	108.4
N1—C1—H1	122.8	H12A—C12—H12B	107.5
S1—C1—H1	122.8	N7—C13—S8	113.8 (3)
N2—C2—S2	127.1 (2)	N7—C13—S7	126.1 (3)
N2—C2—S1	114.6 (2)	S8—C13—S7	120.0 (2)
S2—C2—S1	118.30 (18)	N8—C14—S8	114.9 (3)
C4—C3—S2	106.2 (2)	N8—C14—H14	122.5
C4—C3—H3A	110.5	S8—C14—H14	122.5
S2—C3—H3A	110.5		
N5 ⁱ —Cu1—N1—C1	-137.6 (3)	S2—C3—C4—C5	175.9 (3)
N5—Cu1—N1—C1	42.4 (3)	C3—C4—C5—S3	163.0 (3)

N4 ⁱⁱ —Cu1—N1—C1	-50.3 (3)	C6—S3—C5—C4	74.5 (3)
N4 ⁱⁱⁱ —Cu1—N1—C1	129.7 (3)	N4—N3—C6—S4	0.6 (3)
N5 ⁱ —Cu1—N1—N2	36.6 (2)	N4—N3—C6—S3	178.9 (2)
N5—Cu1—N1—N2	-143.4 (2)	C7—S4—C6—N3	-0.2 (3)
N4 ⁱⁱ —Cu1—N1—N2	123.9 (2)	C7—S4—C6—S3	-178.5 (2)
N4 ⁱⁱⁱ —Cu1—N1—N2	-56.1 (2)	C5—S3—C6—N3	12.7 (3)
C1—N1—N2—C2	0.6 (4)	C5—S3—C6—S4	-169.1 (2)
Cu1—N1—N2—C2	-174.2 (2)	N3—N4—C7—S4	0.7 (4)
C6—N3—N4—C7	-0.8 (4)	Cu1 ^{iv} —N4—C7—S4	-151.87 (18)
C6—N3—N4—Cu1 ^{iv}	158.3 (2)	C6—S4—C7—N4	-0.3 (3)
N1 ⁱ —Cu1—N5—C8	65.1 (3)	N6—N5—C8—S5	-0.5 (3)
N1—Cu1—N5—C8	-114.9 (3)	Cu1—N5—C8—S5	173.93 (14)
N4 ⁱⁱ —Cu1—N5—C8	-26.5 (3)	C9—S5—C8—N5	0.4 (3)
N4 ⁱⁱⁱ —Cu1—N5—C8	153.5 (3)	N5—N6—C9—S5	-0.1 (3)
N1 ⁱ —Cu1—N5—N6	-120.6 (2)	N5—N6—C9—S6	178.9 (2)
N1—Cu1—N5—N6	59.4 (2)	C8—S5—C9—N6	-0.1 (3)
N4 ⁱⁱ —Cu1—N5—N6	147.8 (2)	C8—S5—C9—S6	-179.1 (2)
N4 ⁱⁱⁱ —Cu1—N5—N6	-32.2 (2)	C10—S6—C9—N6	-78.7 (3)
C8—N5—N6—C9	0.4 (4)	C10—S6—C9—S5	100.2 (2)
Cu1—N5—N6—C9	-174.7 (2)	C9—S6—C10—C11	69.0 (3)
C13—N7—N8—C14	0.8 (6)	S6—C10—C11—C12	66.9 (4)
N2—N1—C1—S1	-0.6 (3)	C10—C11—C12—S7	-171.5 (3)
Cu1—N1—C1—S1	174.06 (14)	C13—S7—C12—C11	-83.7 (4)
C2—S1—C1—N1	0.3 (3)	N8—N7—C13—S8	-0.6 (5)
N1—N2—C2—S2	-179.2 (2)	N8—N7—C13—S7	-179.1 (3)
N1—N2—C2—S1	-0.4 (3)	C14—S8—C13—N7	0.2 (4)
C3—S2—C2—N2	1.0 (4)	C14—S8—C13—S7	178.8 (3)
C3—S2—C2—S1	-177.8 (2)	C12—S7—C13—N7	11.3 (5)
C1—S1—C2—N2	0.1 (3)	C12—S7—C13—S8	-167.1 (3)
C1—S1—C2—S2	179.0 (2)	N7—N8—C14—S8	-0.6 (5)
C2—S2—C3—C4	166.7 (3)	C13—S8—C14—N8	0.2 (4)

Symmetry codes: (i) $-x+1, -y, -z+1$; (ii) $-x+1, -y+1, -z+1$; (iii) $x, y-1, z$; (iv) $x, y+1, z$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C3—H3B \cdots O3 ^v	0.97	2.47	3.357 (6)	153
C7—H7 \cdots O3 ^{vi}	0.93	2.51	3.172 (6)	128
C8—H8 \cdots O4 ⁱⁱⁱ	0.93	2.47	3.010 (6)	117
C10—H10A \cdots O2 ^{vii}	0.97	2.50	3.423 (7)	159
C14—H14 \cdots O1 ^{viii}	0.93	2.51	3.419 (7)	167

Symmetry codes: (v) $x-1, y, z$; (vi) $x-1, y+1, z$; (iii) $x, y-1, z$; (vii) $-x+1, -y, -z+2$; (viii) $-x+1, -y+1, -z+2$.

Fig. 1

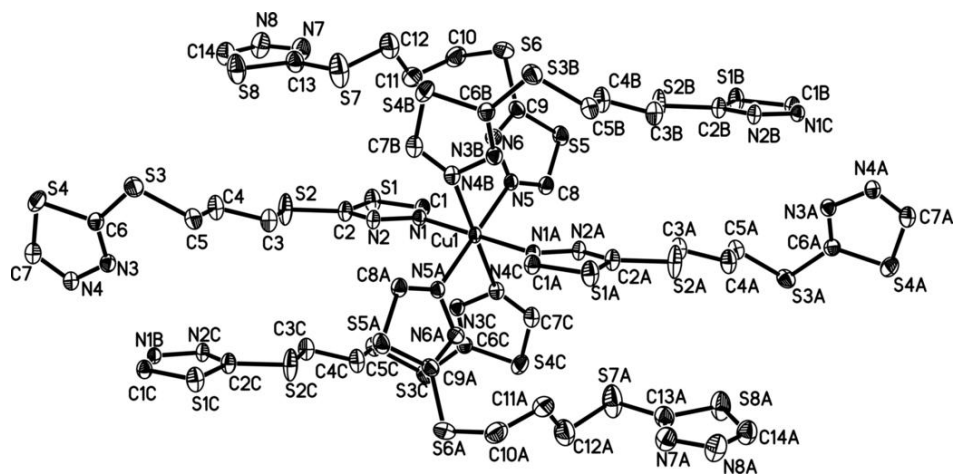


Fig. 2

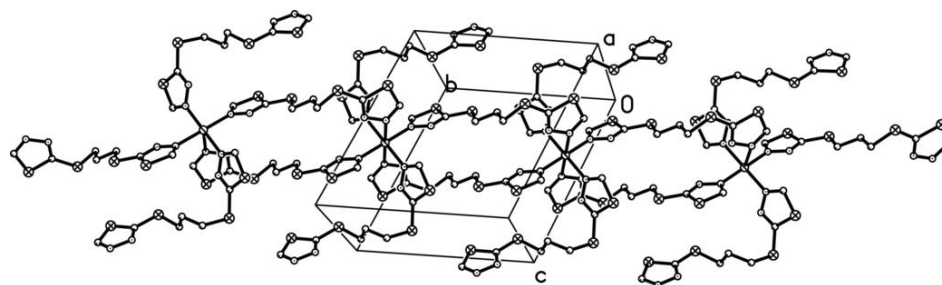


Fig. 3

